

# Photocatalytic degradation of organic pollutants using *Trianthema Portulacastrum* leaf extract based CeO<sub>2</sub> nanoparticles

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#### Abstract:

Comparison of bio CeO<sub>2</sub>-Nps prepared using *Trianthema Portulacastrum* leaf extract with chemical CeO<sub>2</sub>-Nps is of interest. The ultraviolet - visible, x-ray diffraction, HR - TEM, FT - IR, and photoluminescence studies were conducted with CeO<sub>2</sub>-Nps. UV-Maximum absorption at 292 nm was completed using UV-visible spectrum. The HR-TEM images showed 38 nm bio CeO<sub>2</sub>-Nps with spherical morphology. This showed the polycrystalline character of CeO<sub>2</sub>-Nps similar to XRD data. The presence of metal oxide is confirmed by FT - IR analyses. The CeO<sub>2</sub>-Nps showed the potential photocatalytic activity for Acid black 1 color degradation after exposure to sunlight. Chem and bio CeO<sub>2</sub>-Nps have a degradation rate of 86.66 and 94.33%, respectively for acid black 1 dye. The synthesized CeO<sub>2</sub>-Nps are also evaluated for antibacterial and antioxidant activity. The bio CeO<sub>2</sub>-Nps has antibacterial activity for *Pseudomonas aeruginosa* (17 ± 0.56 mm) and *Staphylococcus aureus* (16 ± 0.24 mm) at low concentrations of 100 µl. The CeO<sub>2</sub>-Nps bio showed high inhibition of radical DPPH IC<sub>50</sub> µg/ml, at 95.17 ± 21. Thus, we show that CeO<sub>2</sub>-Nps have environmentally friendly properties that are useful for dye degradation with antimicrobial and antioxidant activities.

**Keywords:** CeO<sub>2</sub> nanoparticles, plant extract, dye degradation, antibacterial, antioxidant

**Background:**

The development of green chemistry to synthesize metal-based nanoparticles with extracts of different plants is gaining momentum in recent years [1]. Environmental impacts by bio nanoparticles are highly commended [2]. The plants have various types of phenolic and flavonoid compounds that help in nanoparticulate formation [3, 4]. It has been found that the extracts from various plants, such as *Cataranthus roseus* [6], *Cocos nucifera* [7], *Beta vulgaris* [8], *Catunareg amspinoso* [9] and *Cyphomandra betacea* were used to synthesize non-toxic nanomaterials [10]. The release of toxic substances affecting the environment by several industrial and research activities is evident [11]. Biosynthesized noble metal (Ag, Au... nanoparticles) are used for environmental friendly detoxification and elimination of harmful and deadly materials. [12, 13]. The mechanism of OH in biodegradation is known [14-16]. The formation  $\pi$  complexes as precursors of OH adduct in hydroxylated by-products with  $\gamma$ -radiolysis is described [17-19].

The NPs play an important role in the removal of organic and inorganic contaminants [20-22]. The bio CeO<sub>2</sub>-NPs synthesis is environmentally friendly and non-toxic [23-24]. CeO<sub>2</sub>-NPs are effective alternatives to degrade dyes and other pollutants [25-26]. Therefore, it is of interest to document the photocatalytic degradation of organic pollutants using bio (*Trianthema Portulacastrum* leaf extract) CeO<sub>2</sub> nanoparticles in comparison with the chemical CeO<sub>2</sub> nanoparticles.

**Materials and methods:****Materials:**

Fresh *Trianthema portulacastrum* leaves were collected from chidambaram rural areas, Tamil Nadu, India. Cerium chloride (CeCl<sub>3</sub>) (99.9%) was obtained and used as received by Sigma - Aldrich, Bangalore, India. *Staphylococcus aureus* (ATCC 6538) and *Pseudomonas aeruginosa* (ATCC 9022) were obtained from the microbial culture collection, the Institute of Microbial Technology in Chandigarh, India. Petri plates were selected with a diameter of about 32 cm and a thickness of 2 cm. All other used reagents are of analytical quality.

**Preparation of CeO<sub>2</sub>-Nps using chemical method:**

Cerium chloride (CeCl<sub>3</sub>) was used without further purification as they were received. CeO<sub>2</sub>-Nps were developed using sol-gel processes [27]. 3.72 g cerium chloride salt taken in 10 ml of deionized water, and ammonia was added drop-by-drop until its pH attained 10. The continuous stirring for another two hours until all the precipitation was over. Filters wash and dry the precipitates overnight. The powder was then calcinated at a temperature of 400°C for two hours in an oven.

**Preparation of CeO<sub>2</sub>-Nps using plant extract method:**

10g *Trianthema Portulacastrum* leaf was powdered and mixed with 100 mL of water at 80 °C. The extract of the leaf was filtered with Whatman No. 1. In 100-ml Erlenmeyer, it was preserved for further use at room temperature. 1:2 v/v CeO<sub>2</sub> were prepared using 10 ml CeCl<sub>3</sub> (contains 3.72g) and a 5 ml leaf extract. At a temperature of 85 °C, the mixture was agitated for 4 hours. The yielding of CeO<sub>2</sub>-Nps observed yellowish-brown color. Also, the precipitate was dried for 4 hours at 400 °C.

**Characterization of CeO<sub>2</sub>-Nps**

TEM images of metal oxide nanoparticles were obtained using a transmission electron microscope (PHILIPS CM200 model) at an operating voltage of 20-200kv with resolution: 2.4 Å. XRD spectra were recorded on the X'PERT PRO model X-ray diffractometer from Pan Analytical instruments operated at a voltage of 40 kV and a current of 30 mA with Cu K $\alpha$  radiation. The FT-IR spectra of powdered CeO<sub>2</sub> were mixed with KBr pellets and are recorded in the 4,000-400 cm<sup>-1</sup> range on a Shimadzu FTIR-8400s. To investigate optical responses and compute the bandgap, the synthesized CeO<sub>2</sub>-NP samples have been subject to UV-vision spectroscopy (Shimadzu UV 1650). The energy for the nanoparticles optical band gap is calculated using the Tauc relation based on the absorption spectrum of the nanoparticles:

$$\hbar h\nu = A (h\nu - E_g)^{1/2}$$

Where  $\hbar$  is a coefficient of optical absorption, the photon energy is  $h\nu$ ,  $E_g$  is a bandgap direct, and  $A$  is a constant that is energy-dependent.

Size of the synthesized CeO<sub>2</sub> Nps can be calculated by applying the following equation [28]:

$$D = 0.9\lambda / H \cos \theta \quad (\text{Scherrer equation})$$

If  $D$  is of crystalline size,  $k$  is of a shape factor ( $K=0.9$  in this work),  $\theta$  is of Bragg angle,  $H$  is of full width at half-maximum and  $\lambda$  is of wavelength of X-ray incident. Photoluminescence (PL) behavior was found at room temperature by FLUOROLOG-3.

**Photocatalytic activity:**

The photocatalytic activities of Chem CeO<sub>2</sub>-Nps and plant mediated CeO<sub>2</sub>-Nps were analyzed using the reactions of acid black 1 dye under-stimulated sunlight irradiation. In that experiment, 100 ml of 0.2 g of fine powder catalyst (Chem CeO<sub>2</sub>-Nps and bio CeO<sub>2</sub>-Nps) and  $3 \times 10^{-4}$  M aqueous acid black 1 dye were taken. Photocatalytic measurement time ranged between 0 and 80 minutes. The suspension allowed the adsorption to stir in

the dark for 10 min to achieve the adsorption-desorption balance between the dye and nanoparticles. Subsequently, the suspension was placed under sunlight and read every ten minutes up to 80 minutes.

On the catalyst surface, the proportion of acid black 1 was estimated following the following ratio [29]:

$$\text{Degradation (\%)} = C_0 - C_t / C_0 \times 100$$

where  $C_0$  is the initial absorption and  $C_t$  is the absorption after different intervals of time.

### Antioxidant studies using DPPH method:

1,1-diphenyl-2-picryl hydroxyl radical methods, as reported on Das *et al.* [30], have been tested in *Trianthema Portulastrum* leaf extract, Chem CeO<sub>2</sub>-NP's, and Bio CeO<sub>2</sub>-Nps. Added to 0.1 mM methanol DPPH radical solution in equal volume, the different concentrations of (25/50/100/125/250/500 µg/ml) sample solution were provided. The reaction mixture was incubated for 60 minutes at room temperature. The mixture has been measured for the optical intensity of 517 nm, which provides antioxidant activity. Ascorbic acid was used for the calibration of the resulting activity as standard. The radical scavenging activity (RSA) percentage of the sample was calculated using the following equation:

$$\% \text{ DPPH radical scavenging} = (\text{Absorbance of control} - \text{absorbance of test sample}) / (\text{Absorbance of control}) \times 100$$

### Antibacterial activity:

Antibacterial properties of fresh leaf extract and prepared nanoparticles biological and chemical method has explored by using disc diffusion technic [31]. It has been studied using the clinical isolation of bacterial cultures Gram-positive bacteria and Gram-negative bacteria *Staphylococcus aureus* and *Pseudomonas aeruginosa* as well. Dissolved nutrient agar was swept into the bacterial suspension, poured through sterile swabs of cotton, and produced with the help of an adjustable cork borer made from stainless steel. At 35°C for 48 hours, the plates were incubated. Ciprofloxacin is used as a positive control and the 50 µl and 100 µl *Trianthema Portulastrum* leaf extracts, the Chem CeO<sub>2</sub>-Nps and bioCeO<sub>2</sub>-Nps were added. Table 3 shows the inhibition zone in diameter (mm).

### Statistical analysis:

The results were evaluated statistically by sigma plot 12.5; an average value for three different replications and a standard error (SE) was determined.

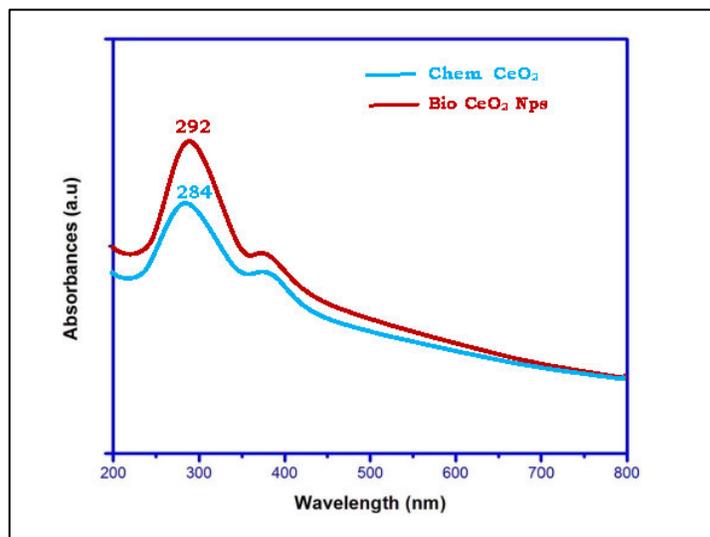


Figure 1: UV-visible spectra of Chem CeO<sub>2</sub>-NPs and biosynthesized CeO<sub>2</sub>-NPs

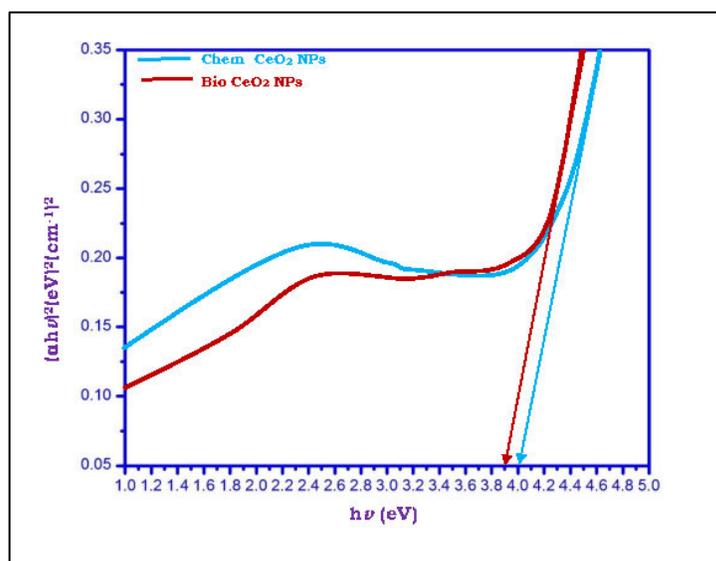
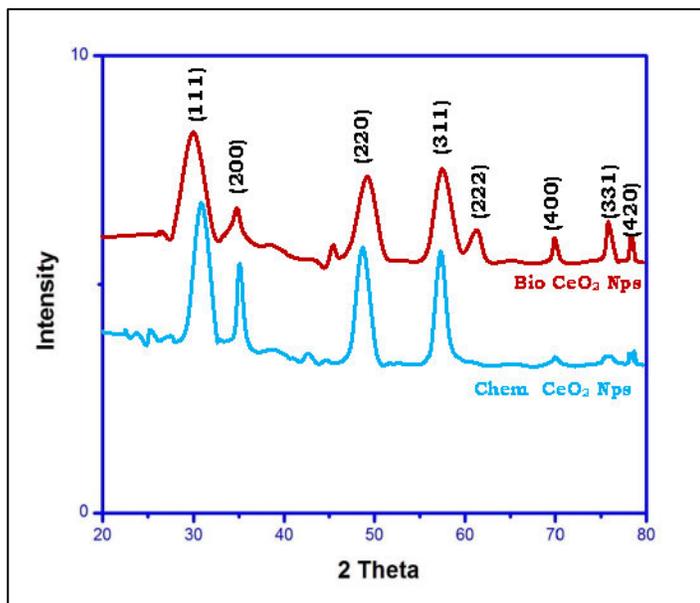


Figure 2: Band gap energy Chem CeO<sub>2</sub>-NPs and biosynthesized CeO<sub>2</sub>-NP

### Results and Discussion:

Chemical CeO<sub>2</sub>-Nps and bio CeO<sub>2</sub>-Nps are measured using the optical absorption (Figure 1). Chem CeO<sub>2</sub>-Nps and biosynthetic CeO<sub>2</sub>-Nps at 284 and 292 nm absorption peaks are observed and all

these values are red shifts relative to the absorption maximum (284 nm) of the Chem CeO<sub>2</sub>. Qaisar *et al* have shown similar absorption peaks (at 315 nm) for bio CeO<sub>2</sub>-Nps [32]. The trianthema Portulastrum extract is comprised of phytochemicals that serve as a cap and reduction agent and also cause the UV absorption point to shift. The absorption position was suggested to depend on the size and shape of the particle in CeO<sub>2</sub>-NP. The UV - visible absorption potential of the CeO<sub>2</sub>-Nps is correlated with the bandgap energy, differentiating between CeO<sub>2</sub>-Nps in different forms. Tauc's equation is used to compute the gap of synthesized CeO<sub>2</sub>-Nps [33].



**Figure 3:** X-ray diffraction pattern of Chem CeO<sub>2</sub>-NPs and biosynthesized CeO<sub>2</sub>-NPs

Chem CeO<sub>2</sub>-NP and bio CeO<sub>2</sub>-NP band gap energy values have been identified 4.00 and 3.90 eV respectively (**Figure 2**). The band gap of the biologically synthesized CeO<sub>2</sub>-Nps can be seen to be smaller than the Chem CeO<sub>2</sub>-Nps. The powerful interaction between CeO<sub>2</sub> and *Trianthema portulastrum* leaf extract phytochemicals (flavonoids and proteins) allows for a faster process for recombining electrons and has resulted in a reduction in band gap for bio CeO<sub>2</sub> Nps. For CeO<sub>2</sub>-NP-biosynthesized, the observed band gap value 3.90 eV is appropriate for photocatalytic and antibacterial activities, which involve electron-exciting formation.

**Figure 3** shows the patterns of X-ray diffraction of chemicals CeO<sub>2</sub>-Nps and bio CeO<sub>2</sub>-Nps with various concentration of *Trianthema portulastrum* leaf extract. The sharp, intense diffraction peaks show Crystal structure and purity. The cubic structure of the CeO<sub>2</sub>-Nps (Jcpds no: 043-2002) is the most responsive Bragg's Peaks that can be reported with the Miller Index (111), (200), [220], (400), (331) and [422] [34-35]. In determining the average crystallite sample size, the Scherrer formula has been used. **Figure 3** shows a crystal size of 78 nm for Chem CeO<sub>2</sub>-Nps. With an increasing percentage of *Trianthema portulastrum* leaf extract; the crystal size decreases for bio CeO<sub>2</sub>-Nps and is found to be 34 nm. Bio CeO<sub>2</sub>-Nps are observed to have a minimum crystallite size due to their quantum confinement effect.

The FT - IR spectroscopy helps to detect leaf extract bio-molecules attached to the CeO<sub>2</sub> surface. FT - IR spectra for *Trianthema Portulastrum* dried leaf extract, Chem CeO<sub>2</sub>-Nps and CeO<sub>2</sub>-Nps bio are displayed in **Figure 4a-c**. **Figure 4a** shows the peaks and their assignments. FTIR spectroscopy illustrated absorption peaks at 3400, 2928, 1720, 1221 cm<sup>-1</sup> were reproduced in the extract of *Trianthema portulastrum* leaf. The absorption band of O-H stretching vibration appears at 3400 cm<sup>-1</sup>. The absorption bands at 2928 and 1720 cm<sup>-1</sup> is due to aldehydic C-H stretching and C=O vibration, respectively. 1231 cm<sup>-1</sup> is due to C-N stretching vibration. Bio CeO<sub>2</sub>-Nps shows FT - IR peaks at 3260, 2310, 1725, 1512, 1255, 1012 and 788 is due to presence of free O - H attachment [36-37], CH vibration, NH primary amines, CH<sub>2</sub> bond, CH<sub>3</sub> is due group, vinyl group and C -O stretching mode vibration [38]. The leaf extracts contain flavonoids that are potent reducing agents that can reduce cerium chloride heptahydrate salt. These flavonoids act as surfactants and are fixed to the CeO<sub>2</sub>-NP surface, and by electrostatic stabilization, they stabilize CeO<sub>2</sub>-NP's. As a result, *Trianthema portulastrum* leaf extract has a dual function to reduce and stabilize CeO<sub>2</sub>-Nps.

Photoluminescence spectroscopy (PL) usually explores the efficiency of the migration and transmission of charging carriers and also the chance of electron-hole pairs in metal oxide [39]. In this research, Photoluminescence spectrum is used to collect significant evidence about surface defects, oxygen vacancies and surface conditions which may sulphurise the impact of the photocatalytic response. The Chem CeO<sub>2</sub>-Nps and bio CeO<sub>2</sub>-Nps show room temperature PL spectrum in Fig 5. The two samples show similar peak positions but vary in intensity. With increasing leaf extract *Trianthema portulastrum* percentage the PL intensity increases. The synthesized CeO<sub>2</sub>-NP emission spectrum includes three peaks of 385, 443 and 469 nm, which reflecting the near-band emissions one violet and two blue emissions. Excitonic recombination is the result

of the Chem CeO<sub>2</sub>-Nps PL emittance peak at 389 nm. It is due to the transitions of 5d-4f of Ce<sup>3+</sup> from ground state 2D(5d1) to state 2F<sub>5/2</sub>(4f1) [40]. At 443 and 469 nm, the emission peak is related to oxygen vacancies [41-42]. The bio CeO<sub>2</sub>-Nps has a blue - shift at 443 nm and 469 nm compared with the chem CeO<sub>2</sub>-Nps. The blue emission peak lies at 443 nm due to the transition from the oxygen vacancy. The oxygen defects in bio CeO<sub>2</sub>-Nps thus support to connect the photo - induced electron easily in excitons. This shows that the intensity of PL has increased. The enhanced PL shows the intensity of bio CeO<sub>2</sub>-Nps ' good crystalline nature and shows desirable catalytic properties.

The morphological and particulate sizes of the synthesized CeO<sub>2</sub>-Nps are demonstrated by high - resolution transmission electron microscopy (HR-TEM). The figures 6a & 7a show typical TEMs obtained with CeO<sub>2</sub>-Nps prepared using trianthema Portulastrum extract and Chem CeO<sub>2</sub>-Nps. Synthesized CeO<sub>2</sub>-Nps have a morphology of almost cubic nanocrystals. In Figures 6d & 7d the histogram showing the distribution of particle size. The histogram in the bioCeO<sub>2</sub>-Nps and chemicals CeO<sub>2</sub>-Nps is narrower in width and the mean particle size is 38 and 82 nm. The particle size seen in HR-TEM is less than the dynamic light scattering value. The electron (SAED) pattern selected for the area is confirmed with the crystal plane nature of a bio CeO<sub>2</sub>-Nps, with the bright-circulated spots that correspond to the following (1 1 1), (2 2 0) (2 2 1), (2 2 2), (4 0 0), (3 3 1) and (4 2 0). The SAED pattern of crystalline impurities shows no other rings [43-44].

#### Photocatalytic activity:

CeO<sub>2</sub>-Nps are environmentally friendly among many rare earth elements, due to their ecologically based photocatalytic application. Industrial waste contains various types of toxic and organic dyes released into water bodies. It has a major environmental impact. All dyeing agents are organically stable. The colors of acid black 1 dye in both oxidized and reduced shapes are different, so it is picked for the study.

For chemical CeO<sub>2</sub>-Nps and bio CeO<sub>2</sub>-NP, photocatalytic activity is conducted to investigate the degradation of an aqueous acid black 1 dye solution by open-air sunlight. In Figure 8a-d you can see the catalytic degradation of the dye. The spectrum UV -Vis is recorded at different intervals 0, 20, 40, 60 and 80 min, between 200 and 800 nm. If it is acid black 1 dye, the peak UV absorption at 345 and 615 nm indicates that the dark blue of the dye becomes a colorless due to electron transition. The bands of 615 nm show that, owing to the catalytic effectiveness Chem CeO<sub>2</sub>-NP and the bio CeO<sub>2</sub>-N Ps, 86.66 and 94.33% of dye are exactamente 80 minutes degraded (Table 1). When the catalyst is added, the increased reduction rate is

observed. This refers to the potential redox enhancement of the electron movement process between beneficiary and recipient. Bio CeO<sub>2</sub>-Nps act as an effective redox catalyst with an electron relay effect. The size of metal nanoparticles plays a major role in catalytic reductions, while the size of bio CeO<sub>2</sub>-Nps has decreased that promotes reactant adsorption on the catalyst surface and simplifies degradation. This will greatly improve the efficiency of the catalyst by increasing the particle surface area. Table 2 shows reusability efficiency of bio and chem CeO<sub>2</sub>, upto two cycles there is no decrease in percentage degradation of acid black 1 dye

#### Mechanistic pathway of dye degradation:

The various quantities of oxygen vacancies show that photocatalytic results are different. It further suggests that significantly more oxygen vacancies will require quick recombination of electron holes and thus decrease photocatalysis for Chemical CO<sub>2</sub>-Nps [45]. The difference in photocatalytic activity has highly been linked in accordance with concentration errors on the nanoparticles surface [45]. They also showed that surface defects have been increased as the particle sizes decreased and photocatalytic activity increased. The present study shows high photocatalytic activity in bio-synthesized CeO<sub>2</sub>-Nps with the smallest particle size attributable to the high separation capacity of the photo generating chargers, large specific areas, and increased absorption of light. Based on these, the possible photo-degradation of the Acid black 1 dye over the UV-radiated CeO<sub>2</sub>-Nps is shown in Figure 8.

The above reaction stages allow electrons (e<sup>-</sup>) to be excited into the conductivity band (CB) by sunlight when the bi-synthesized CeO<sub>2</sub>-NP is radiated by the same number of holes (h<sup>+</sup>) in the VB. Photo-initiated holes react reasonably with Acid black 1 or attach to the surface H<sub>2</sub>O or OH  $\square$  bound to provide a solid oxidant OH  $\square$  radical species. It is suggested that the produced electron binds to O<sub>2</sub> adsorbed to produce O<sub>2</sub><sup>-</sup>. This means that H<sup>+</sup> produces HO<sub>2</sub><sup>·</sup>, which leads to radical OH<sup>·</sup> from the trapped electron. Therefore, the Acid black 1 dye could be degraded by produced reactive species such as OH<sup>·</sup>, HO<sub>2</sub><sup>·</sup>, and O<sub>2</sub><sup>-</sup>.

#### Kinetic studies:

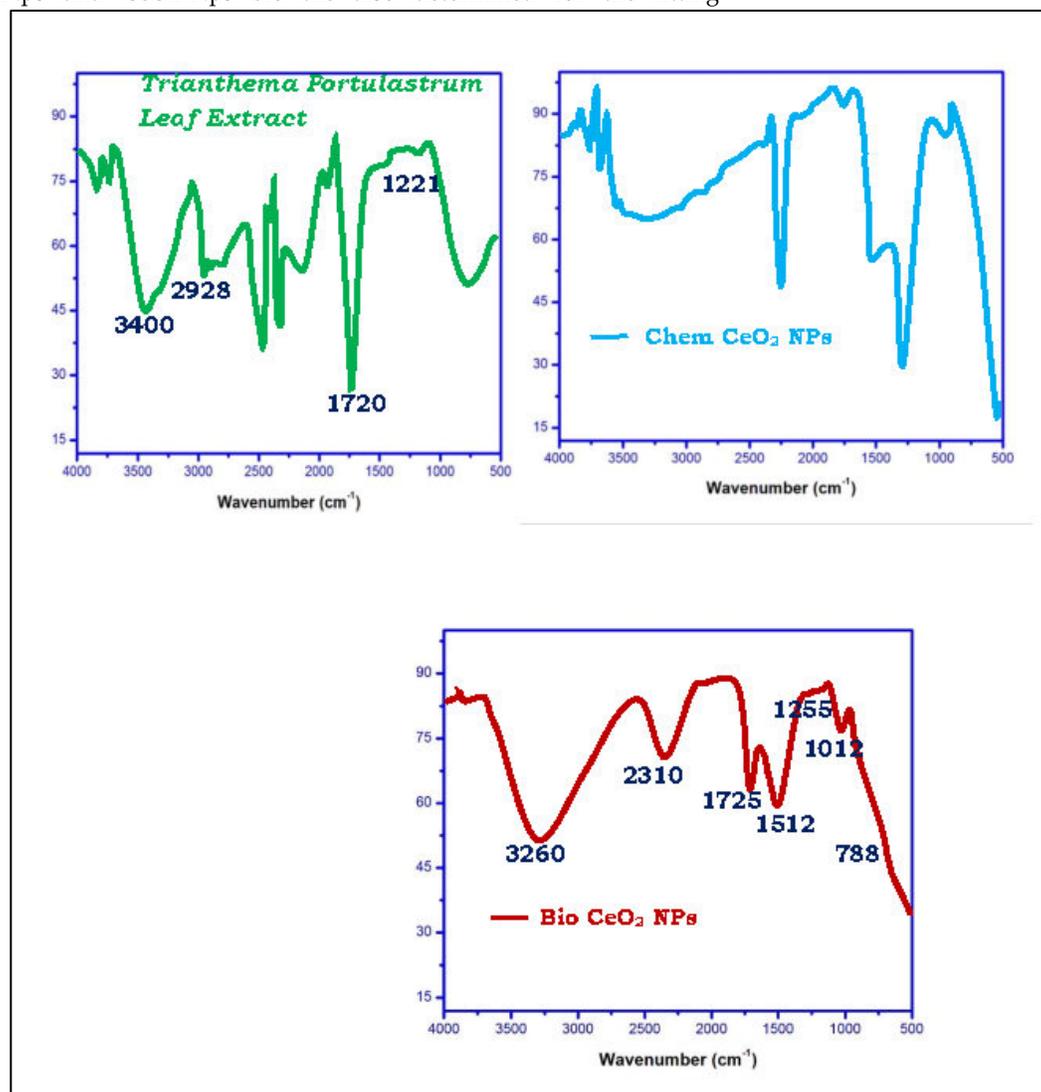
The kinetics of photocatalyst organic degradation in pseudo-first order is described elsewhere [46].

$$\ln(C_0/C_t) = -kt$$

Where k is the apparent reaction rate constant, C<sub>0</sub> is an initial concentration of aqueous Trypan blue, t is a time of reaction and C

is an aqueous Acid black 1 color at a time of reaction  $t$ . Bio-synthesized CeO<sub>2</sub>-Nps and Chem CeO<sub>2</sub>-Nps are investigated and the kinetics of photodegradation of Acid black 1 is presented in **Figure 8a, b**. A pseudo-first-order rate equation determines the rate constant ( $K$ ) for Acid black 1 dye degradation using synthesized CeO<sub>2</sub>-Nps. The graph  $\ln(C_0 / C_t)$  is a rate constant of bio and chemically synthesized CeO<sub>2</sub> Nps 7.8524 and 5.5924 min<sup>-1</sup> based on the irradiation duration. Also, 0.9832 and 0.9750 for Chem CeO<sub>2</sub>-Nps and CeO<sub>2</sub>-Nps bio are also determined for the fitting

correlation coefficient ( $R^2$ ). Finally,  $C_0/C_t$  decreased with time increasing and vice versa. With the increase in time, the percentage of degradation is increased (**Figure 8d**). As a result, BioCeO<sub>2</sub>-Nps demonstrated an improved photocatalytic efficiency in Acid black 1 dye than Chem CeO<sub>2</sub>-Nps and other literature values.



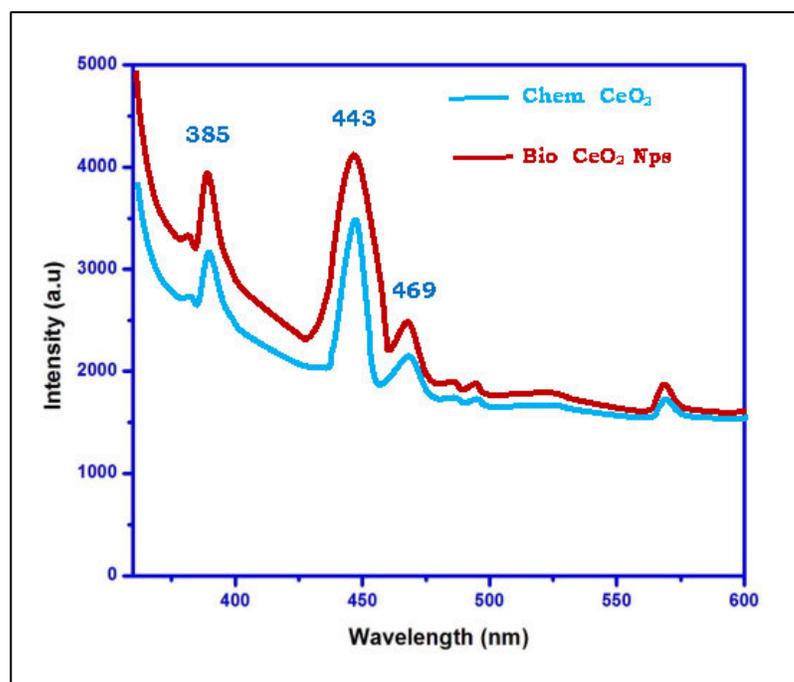
**Figure 4:** FT-IR spectrum of *Trianthema portulastrum* leaf extract, Chem CeO<sub>2</sub>-NPs and biosynthesized CeO<sub>2</sub>-NPs

**Table 1:** % degradation of acid black 1 dye compared to the Chem CeO<sub>2</sub>-NPs and biosynthesized CeO<sub>2</sub>-NPs

Time (min)	% Degradation of acid black 1 dye	
	Bio CeO <sub>2</sub> NPs	Chem CeO <sub>2</sub> NPs
0	0	0
20	32.14	28.41
40	46.58	39.28
60	63.65	54.57
80	75.1	68.74
100	94.33	86.66

**Table 2:** Reusability of acid black 1 dye compared to the Chem CeO<sub>2</sub>-NPs and biosynthesized CeO<sub>2</sub>-NPs

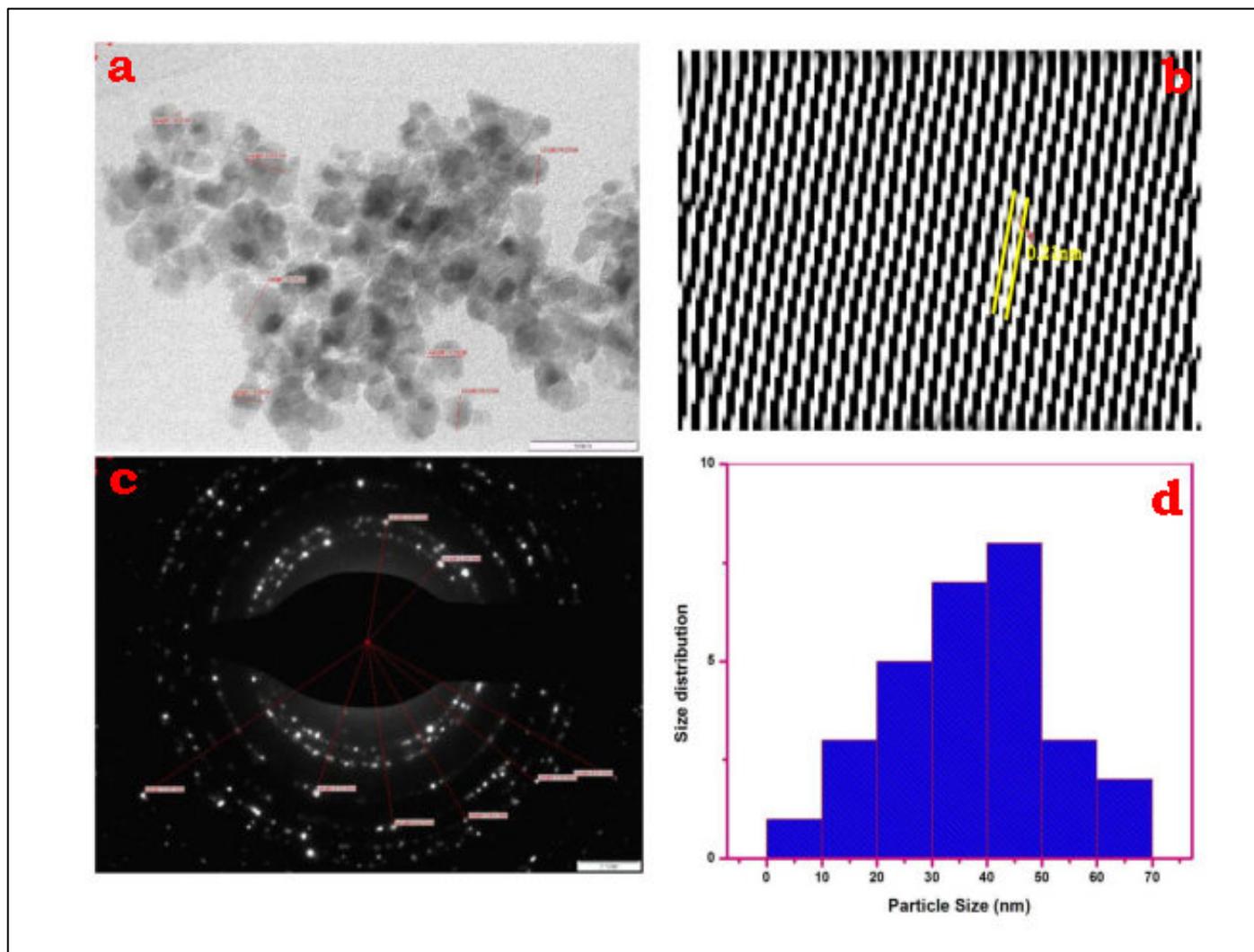
Cycles	1	2	3	4	5
Bio CeO <sub>2</sub>	94	94	92	92	88
Chem CeO <sub>2</sub>	86	86	84	82	80



**Figure 5:** Photoluminescence spectra of Chem CeO<sub>2</sub>-NPs and biosynthesized CeO<sub>2</sub>-NPs

**Table 3:** DPPH free radical assay of Trianthema Portulastrum leaf extract, ChemCeO<sub>2</sub>-NPs and biosynthesized CeO<sub>2</sub>-NPs

Compound	Concentration $\mu\text{g/ml}$						IC <sub>50</sub>
	25	50	100	125	250	500	
Leaf extract	12 $\pm$ 0.23	21 $\pm$ 0.54	35 $\pm$ 0.07	48 $\pm$ 0.28	57 $\pm$ 0.25	69 $\pm$ 0.25	102.52
Chem CeO <sub>2</sub> NPs	21 $\pm$ 0.12	32 $\pm$ 0.87	460.13	57 $\pm$ 0.45	65 $\pm$ 0.14	80 $\pm$ 0.58	104.86
Bio CeO <sub>2</sub> NPs	28 $\pm$ 0.09	39 $\pm$ 0.65	51 $\pm$ 0.45	68 $\pm$ 0.57	76 $\pm$ 0.36	89 $\pm$ 0.47	95.17
Standard	32 $\pm$ 0.45	46 $\pm$ 0.35	59 $\pm$ 0.23	74 $\pm$ 0.31	84 $\pm$ 0.69	96 $\pm$ 0.98	88.49



**Figure 6:** (a) HR-TEM image; (b) lattice fringe; (c) SAED pattern; (d) particle size of biosynthesized CeO<sub>2</sub>-NPs

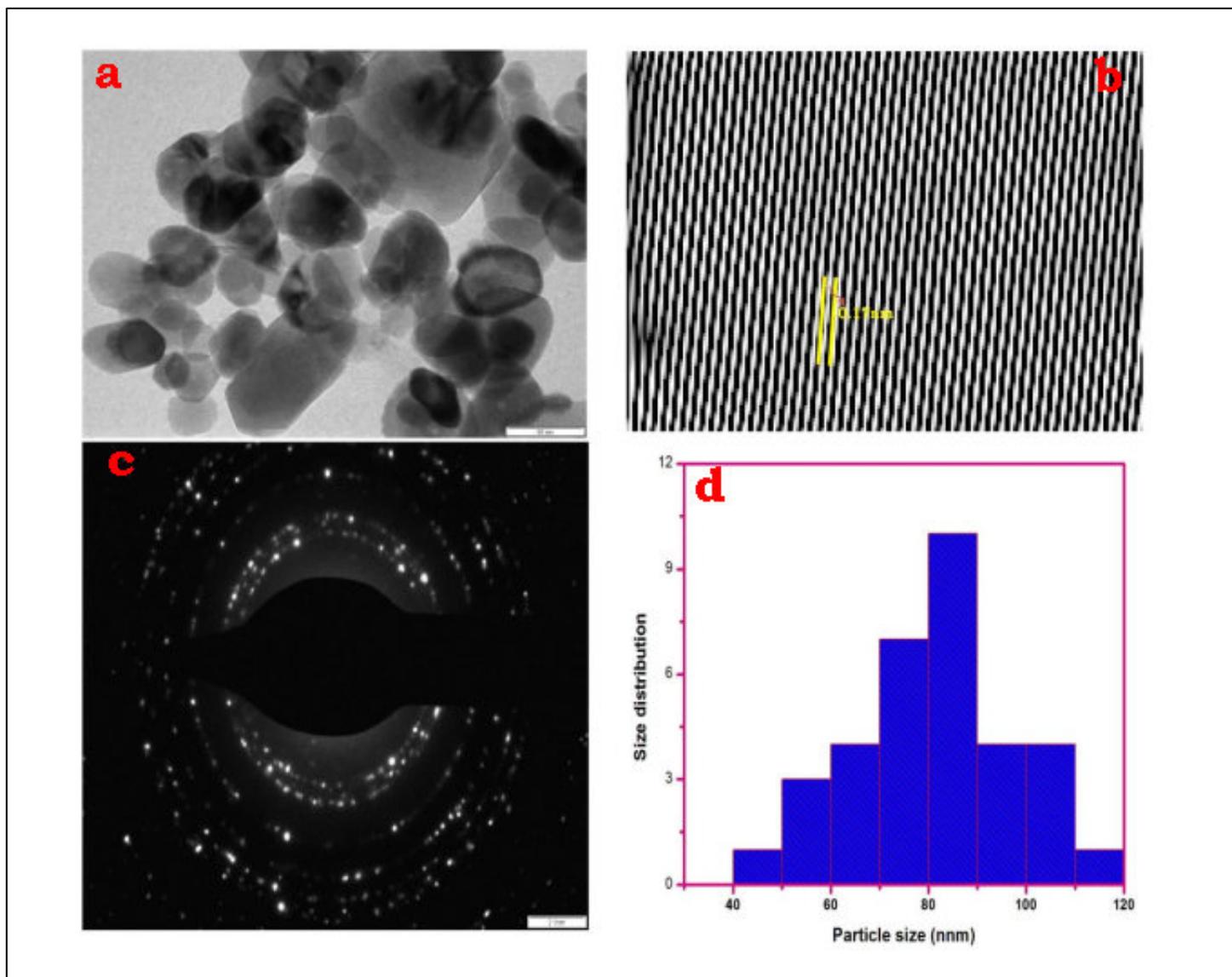
### Antioxidant activity of Synthesized nanoparticles by using DPPH method:

DPPH Radical *Trianthema portulastrum* leaf extract scavenging activity Chem CeO<sub>2</sub>-Nps and CeO<sub>2</sub>-Nps are measured at various concentrations of (25/50/100/125/250/500  $\mu$ g /ml) for standard ascorbic acid. By changing DPPH color, from the initial blue/purple solution to a yellow the reduced activity of *Trianthema portulastrum* leaf extract, bioCeO<sub>2</sub>-NP, and chemCeO<sub>2</sub>-Nps is determined. The percentage of DPPH inhibition is shown in **Figure**

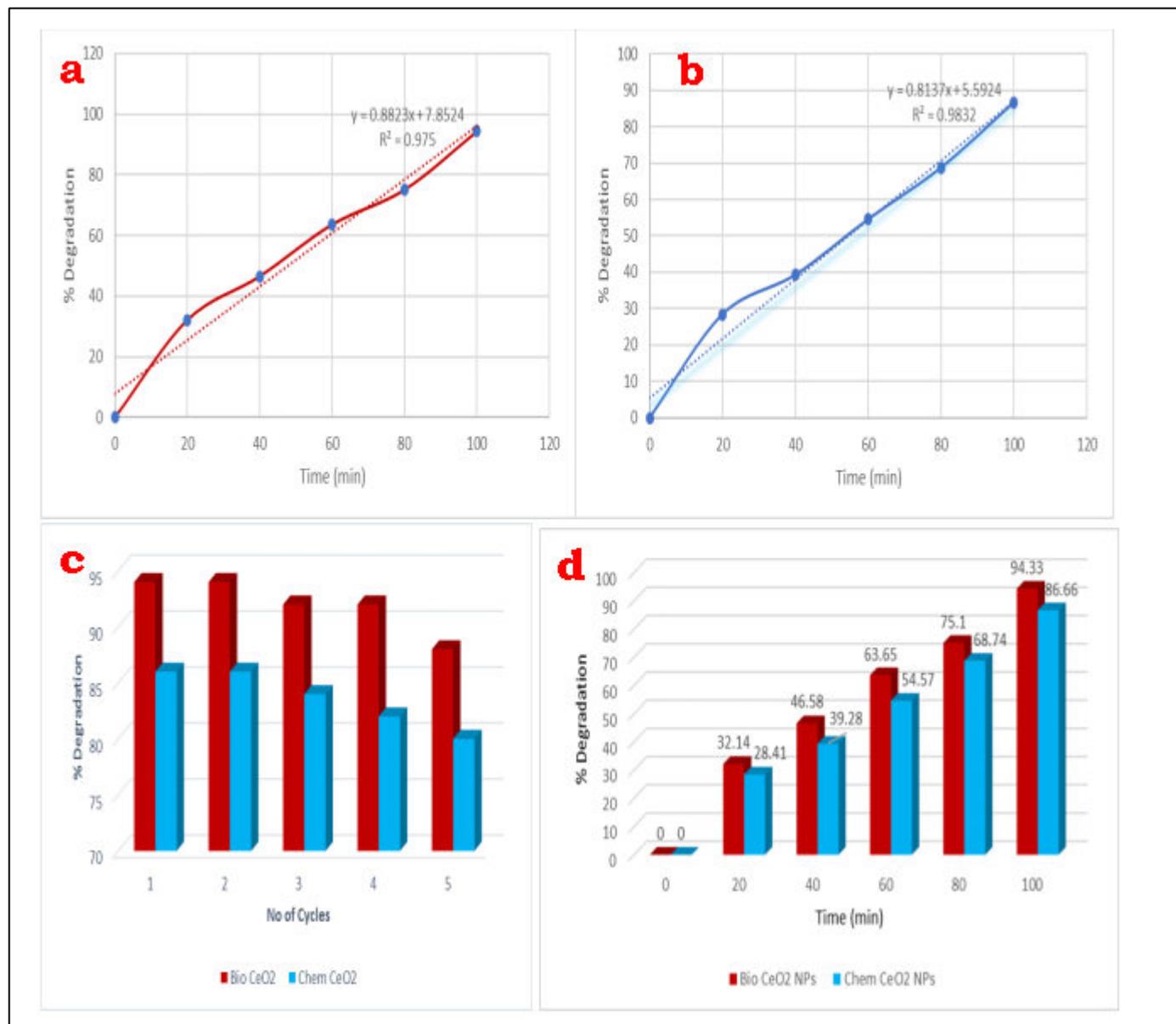
**9 & Table 3.** For *Trianthema portulastrum* leaf extract, chem CeO<sub>2</sub>-Nps bio CeO<sub>2</sub>-Nps, and the standard, the calculated half maximum inhibitory concentration (IC<sub>50</sub>  $\mu$ g / ml) values shall be 102.52, 104.86, 95.17 and 88.49. When IC<sub>50</sub>  $\mu$ g / ml values are lower, the potential for extract antioxidant activity is higher. In comparison to *Trianthema portulastrum* leaf and chemical CeO<sub>2</sub>-NP, the study of DPPH scavenging activity has seen the greatest inhibition in bio CeO<sub>2</sub>-Nps. This result is following Fatemeh et al. studies, which have demonstrated the antioxidant activity of

*Ceratonia siliqua* extract plants using bio CeO<sub>2</sub>-Nps [47]. Moreover, the results of Krishanaveni *et al.* [49] were comparable by the use of *Clitoria ternatea* bio CeO<sub>2</sub>-Nps. Antioxidant activities might be related by the presence of flavonoids, alkaloids in the extract of *Trianthema portulastrum* leaf. This means a reduction in antioxidant activity may result in a reduction in the metabolite concentration of

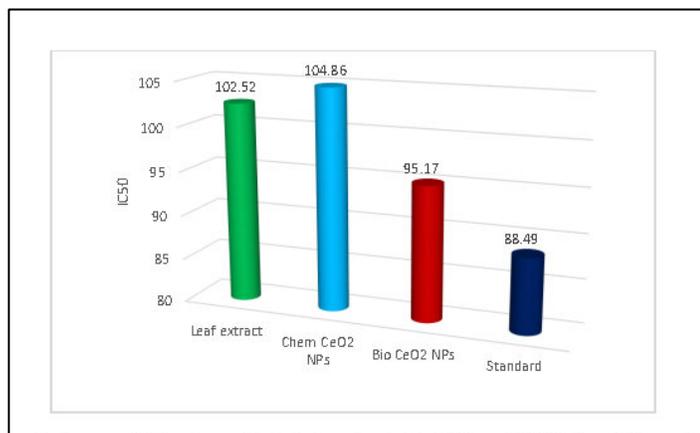
plants during nanoparticulate formation. The surface area of cerium oxide is large, which means more plant chemical substances are added to the active surface. As a result, the shell response phenomenon in the extract of *Trianthema Portulastrum* leaves is elevated by bio CeO<sub>2</sub>-Nps (due to an adsorbed antioxidant moiety on the surface).



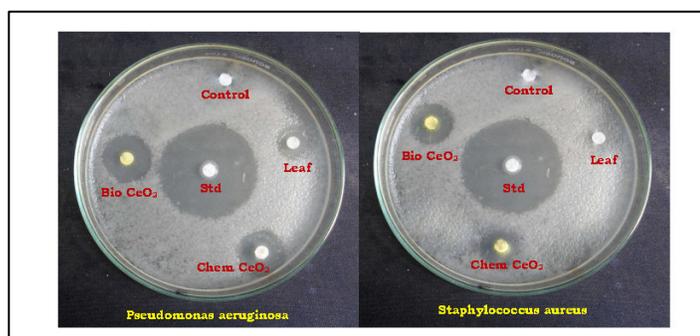
**Figure 7:** (a) HR-TEM image; (b) lattice fringe; (c) SAED pattern; (d) particle size of Chem CeO<sub>2</sub>-NPs



**Figure 8:** (a & b) Rate constant (K) and regression (R<sup>2</sup>); (c) Reusability of biosynthesized CeO<sub>2</sub>-NPs and Chem CeO<sub>2</sub>-NPs; (d) % degradation of acid black 1 dye compared to the biosynthesized CeO<sub>2</sub>-NPs and Chem CeO<sub>2</sub>-NPs



**Figure 9:** DPPH free radical assay of *Trianthema portulastrum* leaf extract, ChemCeO<sub>2</sub>-NPs and biosynthesized CeO<sub>2</sub>-NPs



**Figure 10:** Antibacterial activity of *Trianthema portulastrum* leaf extract, Chem CeO<sub>2</sub>-NPs and biosynthesized CeO<sub>2</sub>-NPs against *Pseudomonas aeruginosa* and *Staphylococcus aureus*

**Table 4:** Antibacterial activity of *Trianthema Portulastrum* leaf extract, Chem CeO<sub>2</sub>-NPs and biosynthesized CeO<sub>2</sub>-NPs against *Pseudomonas aeruginosa* and *Staphylococcus aureus* at 100µL

Compound	Zone of Inhibition (mm)	
	<i>Pseudomonas aeruginosa</i> 100µl	<i>Staphylococcus aureus</i> 100µl
Leaf extract	10±0.08	06±0.41
Chem CeO <sub>2</sub> NPs	14±0.23	11±0.57
Bio CeO <sub>2</sub> NPs	17±0.56	16±0.24
Standard	26±0.89	26±0.11

#### Antibacterial activity by using disc diffusion method:

Bacterial inhibition of *Trianthema portulastrum* extract, ChemceO<sub>2</sub>-NP and bio CeO<sub>2</sub>-Nps are analyzed and the area of inhibition is measured for Gram-positive Bacteria (*Staphylococcus aureus*) and

Grass negative Bacteria (*Pseudomonas aeruginosa*) at 100 µl (Figure 10). Table 4 shows the diameter of the inhibition zone (mm). The bio CeO<sub>2</sub>-Nps (17±0.56 & 16±0.24) exhibit improved bacteriocidal efficiency than Chem CeO<sub>2</sub>-Nps (14±0.23 & 11±0.57) and *Trianthema portulastrum* leaf extracts (10±0.08 & 06±0.41) against *Pseudomonas aeruginosa* and *Staphylococcus aureus*. Particle size and surface area are known to play a key role in their connection with biological cells or to produce secondary damaging products. Due to their size and wide surface area, CeO<sub>2</sub>-Nps produce electronic effects. These electronic effects improve nanoparticles' coupling quality with the microbes CeO<sub>2</sub>-Nps can therefore easily be attached and inserted into the bacteria in the cell membrane [48]. The above mechanisms show that bio CeO<sub>2</sub>-Nps have higher antibacterial activity in comparison with the leaf extract of *Trianthema portulastrum* and Chem CeO<sub>2</sub>-Nps. The increased inhibitory activity of bio ceO<sub>2</sub>-Nps depends not only on the size of nanoparticles and their surface but also on the capping agents (proteins).

#### Conclusion:

The bio and chemical CeO<sub>2</sub>-Nps were synthesized, evaluated, characterized and compared for the photocatalytic degradation of organic pollutants. We show that CeO<sub>2</sub>-Nps degrades acid black 1 coloring under sunlight in a photocatalytic system. Photocatalyst bio CeO<sub>2</sub>-Nps exhibited excellent photocatalytic degradation under visible light irradiation of 94.33%. We also show that the bio CeO<sub>2</sub>-Nps have antibacterial activity. Data show that bio CeO<sub>2</sub>-Nps is associated with various biological and medical applications.

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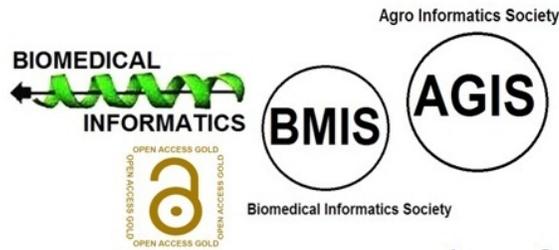
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